

Krystal's Exam #1A (100%) and Bart (100%)



14:57

Back to My eConcordia

1 - What are the coefficients for the reactants and the products when the equation for the reaction:
 $\text{FeCl}_3(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{Fe}(\text{OH})_3(\text{s}) + \text{NaCl}(\text{aq})$

Is balanced?

- 1, 1, 1, 1
 1, 3, 1, 3
 2, 6, 2, 6
 None of the above

2 - What is the mass in grams of 6.022×10^{23} atoms of sulfur?

- 6.022×10^{23} g
 32.06 g
 6.022 g
 1.022×10^{23} g

3 - Sulfur trioxide has the formula:

- SO_3
 S_2O_6
 S_3O
 None of the above

4 - How many liters (expressed in scientific notation and with appropriate number of significant figures) are there in $5.00 \mu\text{L}$.

- 5.00×10^{-6} L
 50.00×10^{-5} L
 (500/1000) L
 0.500×10^6 L

5 - After performing a calculation, Frank's calculator displayed the result as 5.25916. If the answer can have only three significant figures and must be expressed using scientific notation, it should be reported as:

- 5.25
 5.25×10^0
 0.525×10
 5.26×10^0

6 - Element with an atomic number of 54 is a:

- Metal
 Non-metal
 Metalloid
 None of the above

7 - The molar mass of ethanol ($\text{C}_2\text{H}_5\text{OH}$) is given by:

- 92.12 g/mol
- 46.06 g/mol
- 23.03 g/mol
- 46.06

8 - If 5.0 mL of a liquid has a mass of 4.90 g, its density (mass/volume; expressed in scientific notation with appropriate number of significant figures and correct unit) is:

- 0.98 g/mL
- 9.8×10^{-1} g/mL
- 98.0×10^{-2} g mL⁻¹
- 9.80×10^{-1} g/mL

9 - In the Periodic Table, the elements Na, Mg, Al, Si, P, S, Cl and Ar belong to the same:

- Sector
- Period
- Group
- Family

10 - The correct name for H₂O₂ is:

- Dihydrogen dioxide
- Hydrogen peroxide
- Oxygenated water
- Dioxygen dihydride

11 - If an atom contains two protons and two electrons it is:

- A cation
- An anion
- An ion
- An atom with no charge

12 - Anions are formed by the:

- Gain of electrons
- Loss of electrons
- Loss of neutrons
- Gain of protons

13 - An object had a mass of 7.2877 g when weighed on an analytical balance. The uncertain digit in the measurement is:

- 7
- 8
- 2
- None of the above

14 - The final answer for the calculation, $(5.173 \times 10^5) \times (11.5 \times 10^2)$ contains:

- 4 significant figures
- 3 significant figures

- 1 significant figure
- 10^6 significant figures

15 - The smallest volume is represented by:

- 1.0 L
- 1000 mL
- $1.0 \times 10^4 \mu\text{L}$
- $1.0 \times 10^4 \text{cm}^3$

I am ready to submit my answers

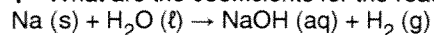
VICTOR'S EXAM #1B (100%)



14:38

Back to My eConcordia

1 - What are the coefficients for the reactants and the products when the equation for the reaction:



is balanced?

- 1, 1, 1, 1
 2, 2, 2, 1
 1, 2, 1, 2
 4, 4, 4, 2

2 - The mass in grams of one mole of carbon atoms is:

- 6.02 g
 12.011 g
 24.04 g
 48.12 g

3 - Ammonium sulfate has the formula:

- NH_4SO_4
 $(\text{NH}_4)_2(\text{SO}_4)_2$
 $(\text{NH}_4)_2\text{SO}_4$
 $\text{NH}_4(\text{SO}_4)_2$

4 - A quantity of 5.44 mg of a substance can be expressed in scientific notation as:

- 5.44×10^0 mg
 5.4×10^0 mg
 (544/100) mg
 0.54×10 mg

5 - After performing a calculation, Peter obtained the result as 7.03961 mL. If the answer can have only three significant figures and must be expressed using scientific notation, it should be reported as:

- 7.04×10^0 mL
 0.703×10 mL
 7.04
 70.4/10 mL

6 - Pure gold is classified as:

- An ion
 An element
 A compound
 A molecule

7 - What is the mass in grams of 6.022×10^{23} atoms of chlorine?

- 6.022×10^{23} g
- 35.453 g
- 6.022 g
- 1.022×10^{23} g

8 - If the density of a substance is 8.140 g/mL, what is the volume in milliliters of a piece of this substance with a mass of 14.99 g?

- 0.526 mL
- 1.899×10^0 mL
- 3.73 mL
- 1.26×10^2 mL

9 - In the Periodic Table, the alkali metals are in the same:

- Group
- Period
- Section
- None of the above

10 - The correct name for $\text{Fe}(\text{OH})_3$ is:

- Iron hydroxide
- Iron trihydroxide
- Iron (III) hydroxide
- Trihydroxide of iron

11 - An atom has 11 protons, 10 electrons and 12 neutrons. The atom is:

- Neutral
- Positively charged
- Negatively charged
- None of the above

12 - A neutral atom containing 20 protons and 40 neutrons loses two electrons. The species formed is:

- An anion
- An ion having a charge of 2+
- Neutral
- An ion having a charge of 2-

13 - Patrick saw 3 cats in his backyard. The number has:

- 1 significant figure
- 0 significant figure
- Infinite number of significant figures
- 3 significant figures

14 - An object had a mass of 14.1763 g when weighed on an analytical balance. The uncertain digit in the measurement is:

- 1
- 8

- 6
- 3

15 - The volume of a room measuring $6.5 \text{ m} \times 6.2 \text{ m} \times 3.0 \text{ m}$ is given by:

- 120.9 m^3
- 120.9
- $1.21 \times 10^2 \text{ m}^3$
- $1.2 \times 10^2 \text{ m}^3$

I am ready to submit my answers

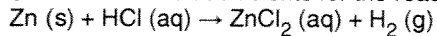
7. Elissa and Jen (100%) Quiz #1C
and Elvina



14:57

Back to My eConcordia

1 - What are the coefficients for the reactants and the products when the equation for the reaction:



is balanced?

- 1, 1, 1, 1
 1, 2, 1, 1
 1, 2, 1, 2
 2, 4, 2, 2

2 - The mass in grams of one mole of silicon atoms is:

- 28.08 g
 56.16 g
 14.04 g
 42.12 g

3 - Calcium sulfate has the formula:

- CaSO_4
 $\text{Ca}_2(\text{SO}_4)_2$
 $\text{Ca}(\text{SO}_4)_2$
 CaS_2O_8

4 - The length of an object is 2.0×10^9 nm. Expressed in meters (in scientific notation with correct number of significant figures) the length is:

- 2.00×10^0 m
 2.0×10^{-18} m
 2.0×10^0 m
 2 m

5 - After performing a calculation, Susan's calculator displayed the result as 5.23961 mL. If the answer can have only three significant figures and must be expressed using scientific notation, it should be reported as:

- 5.24×10^0 mL
 0.524×10 mL
 5.24
 52.4/10 mL

6 - Pure copper is:

- An ion
 An element
 A compound
 A molecule

- 7 - What is the mass in grams of 6.022×10^{23} atoms of carbon?
- 6.022×10^{23} g
 - 12.011 g
 - 6.022 g
 - 1.022×10^{23} g
- 8 - If 5.60 g of a liquid has a volume of 11.2 mL its density is:
- 2.0 mL/g
 - 5.0×10^{-1}
 - 5.00×10^{-1} g/mL
 - 5.0×10^{-1} g/cm³
- 9 - In the Periodic Table, the halogens are in the same:
- Group
 - Period
 - Section
 - None of the above
- 10 - The correct name for Na₂CO₃ is:
- Sodium bicarbonate
 - Sodium carbonate
 - Sodium (II) carbonate
 - Disodium monocarbonate
- 11 - A neutral atom contains 13 protons and has a mass number of 27. The number of neutrons in the atom is:
- 40
 - 13
 - 27
 - 14
- 12 - If an atom has 8 protons and 10 electrons it is:
- A cation
 - An anion
 - Atom with no charge
 - None of the above
- 13 - Which of the following is the SI base unit for length?
- Feet
 - Meter
 - Inch
 - None of the above
- 14 - Object X covered a certain distance in 1.6×10^{-9} s whereas to cover the same distance, object Y took 4.8 ns. Which of following statements is correct:
- X and Y are moving with the same speed

- X is moving at half the speed of Y
- Y is moving faster than X
- X is moving at three times the speed of Y

15 - An object had a mass of 5.3876 g when weighed on an analytical balance. The uncertain digit in the measurement is:

- 7
- 8
- 5
- 6

I am ready to submit my answers



09:56

100% - Exam #2A

Back to My eConcordia

1 - The mere presence of a causative agent constitutes a hazard. It is hazardous in all circumstances.

- True
 False

2 - A diver working at a depth of 8,000 meters may be exposed primarily to which of the following agents?

- Biological
 Physical
 Chemical
 None of the above

3 - The hazard posed by concentrated solutions of potassium permanganate (an oxidizing agent) is classified as:

- Physical
 Chemical
 Biological
 Ergonomic

4 - Which of the following statement is true for acute toxicity?

- Results from brief exposure to high contaminant levels
 Much more difficult to study
 Results from low and repeated exposure over a long period of time
 May have latency periods

5 - In the implementation of WHMIS, suppliers must provide hazard information through:

- Labels only
 MSDSs only
 Labels & MSDSs
 Telephone calls

6 - Chemicals causing changes in DNA are known as:

- Carcinogens
 Mutagens
 Teratogens
 Irritants

7 - Incomplete combustion of oil generates:

- Dust
 Fume
 Smoke
 Gas

8 - The abbreviation TLV stands for:

- Time Limit Value
- Toxic Limit Value
- Threshold Limit Value
- To line up for value

9 - The unit used to express exposure to vapors is:

- m³/L
- L
- ppm
- All of the above

10 - Most common natural routes of entry of chemical agents into the body include:

- Respiratory tract
- Digestive tract
- Skin
- All of the above

I am ready to submit my answers



09:57

100% - Exam #2B

Back to My eConcordia

1 - A causative agent is hazardous only in certain circumstances.

- True
 False

2 - A doctor working in an area with the possibility of exposure to X-rays is at risk of being exposed to which of the following agents?

- Chemical
 Physical
 Biological
 None of the above

3 - The hazard posed by nitric acid (a corrosive and oxidizing material) is classified as:

- Ergonomic
 Biological
 Physical
 Chemical

4 - Which of the following statements is false:

- Severe damage to eyes may result from excessive exposure to ultraviolet radiation
 Microwaves do not penetrate appreciably below the skin
 X-Rays are highly penetrating and damaging to body tissue
 Radiation can be detected by human senses

5 - WHMIS was implemented with the goal to:

- Increase the incidence of illnesses and injuries caused by hazardous materials in the workplace
 Catch those who pollute the environment
 Collect more taxes
 None of the above

6 - Habit forming depressants are classified as:

- Asphyxiants
 Anesthetics
 Narcotics
 Irritants

7 - Extremely fine particles of zinc oxide formed during welding (a high temperature process) are an example of:

- Vapor
 Fume
 Smoke
 Mist

8 - A mountain climber at an altitude of 6000 meters may be subjected to pressure conditions described as:

- Atmospheric
- Hyperbaric
- Hypobaric
- Hyperatmospheric

9 - The abbreviation WHMIS stands for:

- World Human Migration Institute Stand
- Workplace Hazardous Materials Information System
- Workplace Helpful Materials Indexing System
- None of the above

10 - The unit used to express exposure to particles is:

- mg/m^3
- mL
- mg
- All of the above

I am ready to submit my answers



09:57

100% - Exam #2C

Back to My eConcordia

1 - Workplace walk-through surveys are performed to:

- Meet the new workers
- Play cards with the workers
- Pass the newspaper around
- Recognize the potential impact of workplace hazards

2 - A fire fighter trying to put out a fire at a chemical factory is exposed to which of the following agent(s)?

- Chemical only
- Physical only
- Biological only
- Physical and chemical

3 - Ergonomic stresses arise as a result of optimal adjustment between the worker and the work or workplace.

- False
- True

4 - Which of the following statements is false for chronic toxicity?

- Results from brief exposure to high concentrations of the contaminant
- Responses generally have latency periods
- Responses are difficult to observe and relate
- Results from low and repeated exposure over a long period of time

5 - According to WHMIS, the cost of effective worker training must be paid by:

- Employer
- Employee
- Supplier
- Regulators

6 - Chemicals causing malformation in newborns are known as:

- Carcinogens
- Mutagens
- Teratogens
- Irritants

7 - The unit used to express exposure to gases is:

- mL/m³
- ppm
- mL
- All of the above

8 - Science that deals with the properties and interactions of physical, chemical and biological agents is called:

- Sociology
- Toxicology
- Anthropology
- All of the above

9 - The abbreviation MSDS stands for:

- Mechanical Safety Data Sheet
- Material Safety Data Sheet
- Make Sure Dollars Are Safe
- None of the above

10 - The effects of high intensity noise may cause:

- Feeling of hunger
- Interference with communication
- Loss of weight
- None of the above

I am ready to submit my answers

Back to My eConcordia

1 - A concentration of 51.0 ppb for a gas in air is equal to:

- 5.10×10^2 ppb
 5.10×10^{-2} ppm
 5.10×10^5 %
 5.10×10^{-3} %

2 - Based upon TLV-TWA values (given in parentheses), which of the following poses the most risk:

- Compound A (2 ppb)
 Compound B (2.5×10^4 ppb)
 Compound C (100 ppb)
 Compound D (3.0×10^{-4} %)

3 - Capillary action in water makes water in soil available to plants.

- True
 False

4 - The presence of hydrogen bonding in water accounts for its:

- Exceptionally high boiling point
 Taste
 Odor
 Color

5 - The production of ground level ozone is due to:

- Certain photochemical reactions
 Reaction of nitrogen with sunlight
 Reaction of water with oxygen
 All of the above

6 - Nitrogen is a minor component of the air we breathe.

- True
 False

7 - The concentration of air contaminants is determined by:

- Exposing humans to the contaminant
 Sampling and analyzing the contaminant
 Exposing animals to the contaminant
 None of the above

8 - The organization responsible for developing TLVs is abbreviated as:

- ACGIH
 AGCIH
 AHGIC
 ACGHI

9 - A worker is repeatedly exposed to a concentration of 2.5×10^{-4} % of a substance. If the TLV-TWA for the substance is 5.0 ppm, is it a safe work environment for the worker?

- Yes
 No

10 - Hydrogen bonds in water are:

- Weaker than covalent bonds
 Stronger than covalent bonds
 Of the same bond strength
 None of the above

[Back to My eConcordia](#)

1 - A concentration of 21.0 ppm for a gas in air is equal to:

- $2.1 \times 10^{-3} \%$
 $2.10 \times 10^{-6} \%$
 $2.1 \times 10^5 \%$
 $2.10 \times 10^{-3} \%$

2 - Which of the following statements is true based upon the information that the TLVs for two chemicals X and Y are 25 ppm and 45 ppm respectively?

- X poses less risk than Y
 Y poses more risk than X
 X and Y pose the similar risk
 X poses more risk than Y

3 - Oxides of nitrogen in the atmosphere originate both from natural as well as human-made sources.

- True
 False

4 - The nature of O-H bonds in water is:

- Ionic
 Covalent
 Nonionic
 Complex

5 - Ozone containing layer of the atmosphere is called:

- Stratosphere
 Sphere
 Geosphere
 Hydrosphere

6 - Risk characterization integrates the exposure and effects components to estimate the risk.

- True
 False

7 - Air pollutants are responsible for:

- Poor air quality
 Improved air quality
 Cooling the air
 None of the above

8 - SO_x represents:

- SO_3 only
 Mixture of SO_2 and SO_3
 Droplets of H_2SO_4
 Particles of sulfate salts

9 - A worker is repeatedly exposed to a concentration of 2.50×10^3 ppb of a substance. If the TLV-TWA for the substance is 25 ppm, is it a safe work environment for the worker?

- Yes
 No

10 - The attractive forces between water molecules are called:

- Intermolecular
 Nonmolecular
 Intramolecular
 None of the above

I am ready to submit my answers

Back to My eConcordia

1 - If the concentration of SO_2 in air near a pulp and paper mill is 1.5 ppm, what will it be in ppb?

- 1.50×10^3 ppb
- 1.5×10^{-3} ppb
- 1.5×10^3 ppb
- 1.5×10^{-9} ppb

2 - Based upon TLV-TWA values (given in parentheses), which of the following poses the most risk:

- Compound A (2.0×10^7 ppb)
- Compound B (2.5×10^4 ppm)
- Compound C (100 ppm)
- Compound D (3.0 %)

3 - Water is a powerful solvent due to its high polarity and hydrogen bonding ability.

- True
- False

4 - The electron pair in the O-H bond of water is more strongly attracted by the oxygen atom due to its higher:

- Electronegativity
- Molar mass
- Atomic radius
- None of the above

5 - Exposure to O_3 may cause:

- Reduced lung function
- Enhanced lung function
- Hair loss
- Memory loss

6 - Nitrogen is a major component of the air we breathe.

- True
- False

7 - The formation of nitric acid in the atmosphere is due to the oxidation of:

- NO_x
- SO_x
- CO
- All of the above

8 - NO_x represents:

- NO_2 only
- Mixture of oxides of nitrogen
- Droplets of HNO_3
- Particles of nitrate salts



06:15

Exam #4 A

Back to My eConcordia

1 - Adequate safety precautions must be taken in using certain household cleaners (such as oven cleaners) because these may contain:

- Water vapor
- Reactive compounds
- Liquid droplets
- Sharp objects

2 - Non-ionic surfactants are:

- Positively charged
- Negatively charged
- Neutral
- None of the above

3 - In the process of hair coloring, the desired color is obtained by:

- Oxidation of natural hair pigment
- Application of an organic synthetic dye
- Setting the hair in the desired shape
- Treating the hair with a reducing agent

4 - Fragrances in perfumes are generally experienced in three stages or notes because of the:

- Volatility of the ingredients
- Density of the ingredients
- Price of the ingredients
- Color of the ingredients

5 - Sunscreen creams are formulated with additional ingredients to:

- Charge a higher price
- Absorb certain harmful radiation
- Make them look colorful
- None of the above

6 - Skin moisturizers are used to:

- Color the skin
- Maintain the moisture content of the skin
- Dry the skin
- Protect it against sharp objects

7 - Pigments are added to lipstick to act as:

- Deodorants
- Surfactants
- Taste boosters
- None of the above

8 - Cosmeceuticals are believed to be more than simple cosmetics.

- False
- True

9 - One of the influencing factors in making the chemistry of Personal Care Products an expanding and dynamic field is:

- Lower price
- Aging population
- Better packaging of these products
- Bulk production

10 - Hardness of water is due to the presence of which ions?

- Na^+ and K^+
- F^- and Cl^-
- Al^{3+} and Br^-



08:05

Exam #4B

Back to My eConcordia

1 - One of the influencing factors in making the chemistry of Personal Care Products an expanding and dynamic field is:

- Lower price
- Aging population
- Better packaging of these products
- Bulk production

2 - A cationic surfactants is:

- Positively charged
- Negatively charged
- Both positively and negatively charged
- None of the above

3 - Final step in hair coloring process involves using a:

- Dye
- Reducing agent
- Bleaching agent
- Conditioning agent

4 - In creating perfumes, the ingredients are blended based upon their:

- Molecular size only
- Molecular size and volatility
- Volatility only
- Color

5 - Sunscreen products contain chemicals to absorb:

- Microwaves
- Vibration
- Infrared radiation
- UV Radiation

6 - The effective chemical ingredients in deodorants are:

- Antibacterial agents and perfumes
- Water and table salt
- Pigments
- Dyes

7 - Skin moisturizers are used to:

- Color the skin
- Dry the skin
- Protect it against sharp objects
- None of the above

8 - Laundry detergents are formulated by mixing many ingredients because they perform diverse functions.

- True
- False

9 - The impact of household cleaners on the environment is minimized because these products are formulated:

- To work with water and end up in wastewater treatment plants
- Free of chemicals
- With natural ingredients only
- To change into environmentally friendly ingredients immediately after use

10 - Hardness of water is due to the presence of which cations?

- Ca^{2+} and Mg^{2+}
- Na^+ and K^+
- F^- and Cl^-
- Al^{3+} and Br^-



08:12

Exam 14C

Back to My eConcordia

1 - Surface tension of water must be reduced in order for water to:

- Wet a surface effectively
- Freeze easily
- Become denser
- Dissolve more oxygen

2 - Surfactants are effective in:

- Reducing the density of water
- Increasing the surface tension of water
- Decreasing the surface tension of water molecules
- Heating water molecules

3 - Hair conditioners generally contain:

- Anionic surfactant
- No surfactant
- Cationic surfactant
- Water only

4 - Fragrances for perfumes are created by using:

- Naturally occurring compounds only
- Synthetic compounds only
- Both natural and synthetic compounds
- None of the above

5 - Sunscreen creams are formulated with additional ingredients to:

- Charge a higher price
- Absorb infrared radiation
- Make them look colorful
- None of the above

6 - The effective chemical ingredients in deodorants are antibacterial agents to:

- Culture bacteria
- Destroy bacteria responsible for converting certain compounds to unpleasant odors
- Mask the unpleasant odors
- Trap the unpleasant odors

7 - Pigments are added to lipstick to act as:

- Deodorants
- Surfactants
- Colorants
- Taste boosters

8 - An anionic surfactant has a negative charge.

- True
- False

9 - Adequate safety precautions must be taken in using certain household cleaners (such as oven cleaners) because these may contain:

- Water vapor
- Reactive compounds
- Liquid droplets
- Sharp objects

10 - The oxidizing agent generally used during curling of hair is:

- H₂O
- H₂O₂
- O₃

1 - Carbon can form single, double and triple bonds with other elements including carbon.

- True
- False

2 - The compound with the formula ROR' is:

- An amine
- A carboxylic acid
- An ether
- None of the above

3 - The compound with the formula C_2H_5OH is an:

- Alcohol
- Ether
- Amine
- Ester

4 - How many hydrogen atoms are there in a compound with the formula $CH_3(CH_2)_3NH_2$?

- 4
- 9
- 11
- 8

5 - Cephalosporins are used as:

- Antiviral drugs
- Antibiotics
- Anti-inflammatory drugs
- None of the above

6 - SSRIs are used as:

- Fungicide
- Insecticide
- Growth hormone
- None of the above

7 - The three-drug cocktail treatment (HAART) is used to treat:

- AIDS
- Pneumonia
- Cancer
- Depression

8 - Chemotherapy involves treatment of cancer using:

- Hypnotherapy
- Surgery
- Alkylating agents
- None of the above

9 - Certain steroids are used as anti-inflammatory drugs because these drugs:

- Cost a lot of money
- Are antidepressants
- Emit radiation
- None of the above

10 - The number of carbon atoms in a benzene ring is:

- 6
- 9
- 12
- None of the above



09:55

Exam #6A

[Back to My eConcordia](#)

1 - Weakly acidic nature of rain is due to dissolved:

- O₂ (g)
- O₃ (g)
- N₂ (g)
- None of the above

2 - Gaseous carbon dioxide is generated when limestone reacts with:

- Acid
- Base
- Water
- Oxygen

3 - A water sample shows a pH of 3.2. The sample is:

- Acidic
- Neutral
- Colorless
- Basic

4 - The solubility of lead at a pH of 3.2 is higher than at a pH of 9.7.

- False
- True

5 - Solar energy is an example of a:

- Fossil energy source
- Non-fossil energy source
- No energy source
- Wind energy source

6 - One of the gases responsible for global warming is:

- CO
- N₂
- N₂O
- O₂

7 - Reduction in carbon dioxide emissions is a suggested strategy to control global warming.

- False
- True

8 - The depletion of stratospheric ozone is caused by:

- Natural sources only
- Human-made sources only
- Both natural and human-made sources
- None of the above

9 - The pH value for a solution of sodium hydroxide (a basic solution) should be:

- Above 7.0
- Below 7.0
- Equal to 7.0
- Equal to 1.5

10 - The production of acid rain is due to the reactions of ozone and nitrogen present in the atmosphere.

- True
- False

I am ready to submit my answers



09:52

Exam #6B

Back to My eConcordia

1 - Weakly acidic nature of rain is due to dissolved:

- CO₂ (g)
- O₂ (g)
- O₃ (g)
- N₂ (g)

2 - Which of the following compound reacts with an acid in the rain to generate carbon dioxide gas?

- H₂O
- CaCO₃
- NaCl
- NH₃

3 - The pH of a solution is measured to be 10.5. The solution is:

- Acidic
- Basic
- Neutral
- None of the above

4 - The solubility of aluminum at a pH of 9.5 is higher than at a pH of 4.2.

- False
- True

5 - Hydroelectric is an example of a:

- Fossil energy source
- Non-fossil energy source
- No energy source
- Solar energy source

6 - The gas responsible for global warming is:

- CO
- N₂
- O₂
- None of the above

7 - Using non-fossil energy sources is a suggested way to control global warming.

- True
- False

8 - The agreement to reduce and ultimately ban the manufacture and use of CFCs goes by the name of:

- Kyoto Protocol
- EC Free Trade Agreement
- Geneva Convention
- Montreal Protocol

9 - The pH value for a solution of ammonium chloride (an acidic solution) should be:

- Above 7.0
- Below 7.0
- Equal to 7.0
- None of the above

10 - The welders are at a risk of being exposed to ozone in their profession.

- True
- False

I am ready to submit my answers



09:56

Exam #6C

Back to My eConcordia

1 - The production of acid rain is due to the reactions of:

- N₂
- NO_x and SO_x
- CO
- O₃

2 - The reaction of calcium carbonate with sulfuric acid generates gaseous:

- CO
- CO₂
- O₂
- O₃

3 - The pH of a solution is measured to be 7.0. The solution is:

- Acidic
- Basic
- Neutral
- None of the above

4 - The solubility of lead at a pH of 9.5 is higher than at a pH of 4.2.

- True
- False

5 - Nuclear energy is an example of a:

- Fossil energy source
- No energy source
- Wind energy source
- None of the above

6 - One of the gases responsible for global warming is:

- CO
- N₂
- O₂
- CO₂

7 - Energy conservation is one of the suggested ways to control global warming.

- False
- True

8 - In the stratosphere, oxygen and ozone interact with solar energy to form a:

- Dynamic system
- Static system
- Solar system
- No system

9 - The pH value for a solution of sodium chloride (a neutral solution) should be:

- Above 7.0
- Below 7.0
- Equal to 7.0
- Equal to 11.5

10 - Interaction of high intensity UV light with atmospheric oxygen leads to the formation of ozone.

- False
- True

I am ready to submit my answers