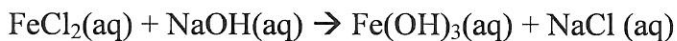


# EXAM 1

1) What are the coefficients for the reactants and the products when the following reaction is balanced:



- a) 1, 1, 1, 1
- b) 1, 3, 1, 3
- c) 2, 6, 2, 6
- d) None of the above

*Never balances*

2) In the balanced equation for the reaction between gaseous nitrogen and hydrogen gas forming gaseous ammonia, the coefficients for hydrogen gas is:

- a) 4
- b) 1
- c) 2
- d) 3



3) What is the mass of one mole of sulphur atoms?

- a)  $6.022 \times 10^{23}$  g
- b) 16 g
- c)  $1.022 \times 10^{23}$  g
- d) None of the above

*From periodic table  $\rightarrow$  32g*

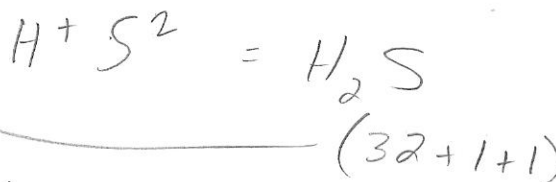
4) One mole of silver atoms contains

- a) 1 atom of silver
- b)  $6.022 \times 10^{23}$  atoms of silver
- c)  $6.022 \times 10^{-23}$  atoms of silver
- d) 1 g of silver

*1 mole =  $6.022 \times 10^{23}$  atoms*

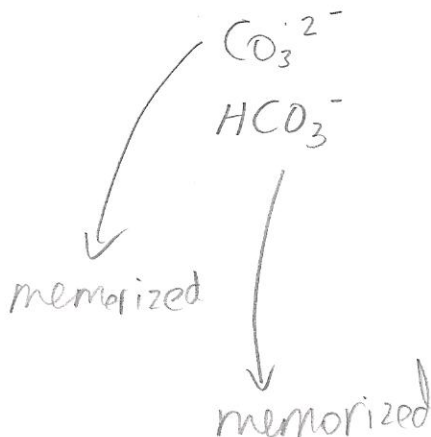
5) The molar mass of hydrogen sulphide is given by

- a) 34.08 g/mol
- b) 1704 g/mol
- c) 68.16 g/mol
- d) 66.06 g/mol



6) The formula for the hydrogen sulphate ion is

- a)  $\text{H}_2\text{SO}_4$
- b)  $\text{HSO}_4^-$
- c)  $\text{HSO}_4^{2-}$
- d)  $\text{SO}_4^{2-}$



*module 1.4  $\rightarrow$  slide 10 table*

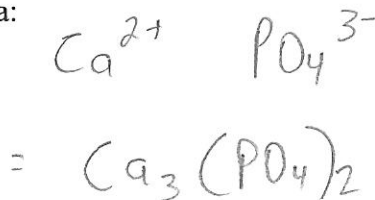
7) The correct name for  $\text{SiF}_6$  is:

- a) 0 Fluoride of silicon
- b) Monosilicon tetrafluoride
- c) Silicon hexafluoride
- d) None of the above

2 non-metals so use hexo, etc  
Never start with "mono"

8) Calcium phosphate has the formula:

- a)  $\text{Ca}_3(\text{PO}_4)_2$
- b)  $\text{Ca}_2(\text{PO}_4)_3$
- c)  $\text{CaPO}_4$
- d)  $\text{Ca}_3(\text{PO}_4)_3$



9) The correct name for  $\text{H}_2\text{O}_2$

- a) Dihydrogen dioxide
- b) Hydrogen peroxide
- c) Oxygenated water
- d) Dioxygen dihydride

(Special name)

10) How many microliters are there in 282L?

- a) 0.282  $\mu\text{L}$
- b) 28.2  $\mu\text{L}$
- c)  $2.82 \times 10^3 \mu\text{L}$
- d)  $2.82 \times 10^8 \mu\text{L}$

micro =  $10^{-6}$

Small to big  $\rightarrow 282 \times 10^6$   
 $2.82 \times 10^8$

11) A mass of 1.7 ng is also equal to

- ~~a)  $1.70 \times 10^{-9} \text{g}$~~
- b)  $1.7 \times 10^9 \text{g}$
- c)  $1.7 \times 10^{-6} \text{mg}$
- ~~d)  $1.70 \times 10^{12} \text{mg}$~~

$$1.7 \times 10^{-9} \text{g}$$
$$1.7 \times 10^{-6} \times 10^{-3} = 1.7 \times 10^{-9} \text{g}$$
$$= 1.7 \times 10^{-6} \text{mg}$$

12) What is the speed (in kilometres per hour) of a car if it takes  $9.02 \times 10^{-1}$  minutes to cover a distance of  $1.00 \times 10^3 \text{m}$ ?

- a)  $1.84 \times 10^{-2} \text{ km/hour}$
- b) 66.5 km/hour
- c)  $1.11 \times 10 \text{ km/hour}$
- d)  $1.11 \times 10^{-1} \text{ km/hour}$



$$\frac{1000 \text{m}}{60 \text{ mins}} = \frac{1 \text{ km}}{1 \text{ hr}}$$
$$9.02 \times 10^{-1} \text{ min} = X$$

$$X = 0.015033$$

$$s = \frac{1}{X} = 66.5 \text{ km/hr}$$

13) The density of a substance is 1.70 g/ml, what is the mass of  $9.20 \times 10^{-3}$  L of this substance?

- a) ~~9.1202 g~~
- b) 5.41 g
- c)  $1.56 \times 10$  g
- d)  $1.85 \times 10^{-1}$  g

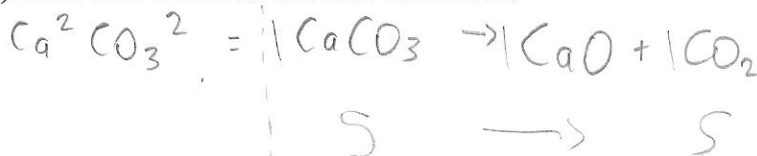


$$m = d \times v$$

$$m = 1.70 \frac{\text{g}}{\text{ml}} \times 9.20 \times 10^{-3} \times \frac{1000}{1}$$

14) If under a set of optimum conditions, calcium carbonate decomposes to produce calcium oxide and carbon dioxide, how many moles of carbon dioxide will be produced (under similar conditions) from 5.00 moles of calcium carbonate?

- a) 5.0 moles
- b) 5.00 moles
- c) 10.00 moles
- d) 2.50 moles



15) A cation  $M^{3+}$  contains 10 electrons. The number for protons from which the cation is formed is given by?

- a) 10
- b) 13
- c) 16
- d) 30



$$\square - 3 = 10$$

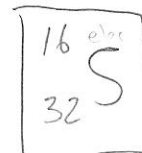
$$\square = 13$$

(+ = you lost electrons)

16) The sulphide ion contains:

- a) 16 protons and 16 electrons
- b) 16 protons and 14 electrons
- c) 14 protons and 16 electrons
- d) 16 protons and 18 electrons

it is negative



17) A neutral atom contains 13 protons and has a mass number of 27. The number of electrons in the atom is

- a) 40
- b) 14
- c) 27
- d) 13

neutral = same electrons & protons

$$(\text{neutrons} = 27 - 13)$$

18) An atom with 11 protons and 12 neutrons represents an atom of

- a) Mg
- b) H
- c) V
- d) Na

check table and look

$$\text{for atomic \# } 11 = \text{Na}$$

19) An element containing 33 protons and 33 electrons is a

- a) Metal
- b) Non-metal
- c) Metalloid
- d) None of the above

look up 33 = As = Metalloid

20) The correct answer (expressed in scientific notation) for the calculation  $229.55\text{g} + 123.345\text{g} + 13.2\text{g}$  is

- a)  $385.105\text{g}$
- b)  $3.661 \times 10^2\text{g}$
- c)  $366.11\text{g}$
- d)  $3.7 \times 10^2\text{g}$

Scientific notation. Answer can only have 1 decimal

21) The mass of an empty container is  $23.0960\text{g}$ . A solid is then placed in the container. The combined mass of the solid and the container is  $23.1190\text{g}$ . How many significant figures are there in the calculated mass of the solid?

- a) 6
- b) 4
- c) 3
- d) 2

$$23.1190 - 23.0960 = 0.0230$$

(there must be 4 decimals)  
leading sig figs don't count

22) The largest volume is represented by:

- ~~a) 1.0L~~
- ~~b) 1000 ml~~
- c)  $1.0 \times 10^5\ \mu\text{L}$
- d)  $1.0 \times 10^4\ \text{cm}^3\ \text{ml}$   
 $\text{cm}^3 = \text{ml}$

$$1.0 \times 10^5 \times 10^{-6} = 10^{-1}\ \text{l}$$
$$1.0 \times 10^4 \times 10^{-3} = 10^1\ \text{l}$$

23) After performing a calculation, Kelly's calculator displayed the result as  $3.45973$  if the answer can only have three significant figures and must be expressed in scientific notation, it should be reported as:

- a)  $3.46 \times 10^0$
- b)  $0.346 \times 10$
- c)  $3.46$
- d)  $3.46/10$

a) is only one in scientific notation

24) In the periodic table, the alkaline earth metals are in the same:

- a) Period
- b) Group
- c) Section
- d) Row

25) Workers have a right to

- a) Work in a safe and healthy environment
- b) Know the processes and substances they are working with
- c) Know the potential hazards from the substance and processes
- d) All of the above

26) According to WHMIS the cost of effective worker training must be paid by

- a) Employee
- b) Supplier
- c) Regulators
- d) None of the above

Answer = employer

27) WHMIS was developed through the collective efforts

- a) Labor, Industry
- b) Federal/Provincial/Territorial Governments
- c) Labor, Provincial Government
- d) Labor, Industry, Federal/Provincial/Territorial Governments

5 things  
↓

28) The mere presence of a causative agent constitutes a hazard, it is a hazard in all circumstances

- a) True
- b) False

29) A mountain climber at an altitude of 5000 meters may be subjected to pressure conditions described as

- a) atmospheric
- b) hyperbaric
- c) hyperatmospheric
- d) hypobaric

hyper = high pressure  
hypot = low pressure

30) Which of the following statements is false?

- a) severe damage to eyes may result from excessive exposure to UV radiation
- b) Microwaves do not penetrate appreciably below the skin
- c) X-rays are not damaging to body tissue
- d) Radiation cannot be detected by human senses

31) Hazard posed by exposure to excessive noise can be classified as

- a) Chemical
- b) Biological
- c) Ergonomic
- d) Physical

32) It has been established that ergonomics, properly applied can help increase job efficiency

- a) True
- b) False

↓  
(furniture)

33) Which of the following statements is true for acute toxicity?

- a) Results from brief exposure to a high concentration
  - b) Responses generally have latency periods
  - c) Responses are difficult to observe and relate
  - d) Results from low and repeated exposure over a long period of time
- Chronic {

34) Chemicals causing malformations in newborns are classified as

- a) Asphyxiants
- b) Carcinogens
- c) Irritants
- d) None of the above

right answer = teratogens

35) The formation of nitric acid in the atmosphere is due to the reactions of

- a)  $\text{NO}_x$
- b)  $\text{SO}_x$
- c) CO
- d)  $\text{SO}_4^{2-}$

↓  
 $\text{NO}_x$

36) Extremely fine particles of zinc oxide formed during welding (a high temperature process) are an example of:

- a) vapour
- b) fume
- c) smoke
- d) mist

} memorize

37) The concentration of air contaminants is determined by

- a) Exposing humans to the contaminants
- b) Exposing animals to the contaminants
- c) Exposing bacteria to the contaminants
- d) Sampling and analyzing the contaminants

38) Commonly used units of concentration for gases are

- a) ppm
- b) ppb
- c) %
- d) None of the above

39) A concentration of 31.0 ppb for a gas is equal to

- a)  $3.1 \times 10^{-3} \%$
- b)  $3.10 \times 10^{-6} \%$
- c)  $3.1 \times 10^5 \%$
- d)  $3.10 \times 10^{-3} \%$

ppb  $\rightarrow$  ppm =  $\div 1000$   
 $\rightarrow$  % =  $\div 10000$

40) The concentration of SO<sub>2</sub> in air near a pulp and paper mill is 23. ppb, what will it be in ppm

- a)  $2.30 \times 10^3$  ppm
- b)  $2.3 \times 10^{-2}$  ppm
- c)  $2.3 \times 10^3$  ppm
- d)  $3 \times 10^{-9}$  ppm

$$\frac{23}{1000} =$$

$$2.3 \times 10^{-2}$$

41) Based upon TLV-TWA values (given in parenthesis) which compound poses the most risk?

- a) Compound A ( $2.0 \times 10^7$  ppb)  $2.0 \times 10^4$
- b) Compound B ( $2.0 \times 10^4$  ppm)
- c) Compound C (100 ppm)
- d) Compound D (3.0 %)  $3.0 \times 10^4$

42) A worker is repeatedly exposed to a concentration of  $9.5 \times 10^3$  ppb of a substance if the TLV TWA for the substance is 2.5 ppm it is a safe work environment for the worker

- a) True
- b) False

$\downarrow$   
 $\div 1000 = 9.5 = \text{not safe}$

43) The organization responsible for developing TLVs is

- a) AGCIH
- b) ACGHI
- c) AHGIC
- d) None of the above

ACGIH

44) Most common natural routes of entry of chemical agents into the body include

- a) respiratory tract
- b) digestive tract
- c) skin
- d) none of the above

inhaling

45) Methanol (a polar covalent compound) dissolves in water by the formation of

- a) Covalent bonds
- b) Ionic bonds
- c) Hydrogen bonds
- d) Oxygen bonds

~~hydrogen bonds~~

46) The presence of hydrogen bonding in water accounts for its

- a) Taste
- b) Purity
- c) Odor
- d) None of the above

many other things

47) Hydrogen bonds in water are:

- a) Weaker than covalent bonds
- b) Stronger than covalent bonds
- c) Of the same bond strength
- d) None of the above

48) Surfactants are effective in

- a) Increasing the density of water
- b) Increasing the surface tension of water
- c) Increasing the polarity of water
- d) None of the above

right answer:  
decreasing surface tension  
of water

49) In the cleaning process the polar end of the surfactant is attracted to water because water is

- a) polar
- b) dirty
- c) non-polar
- d) hot

polar → polar  
non → non

50) An anionic surfactant is

- a) negatively charged
- b) positively charged
- c) without a charge
- d) none of the above

Cationic = positive

51) Detergents almost eliminate the scum formation because they are

- a) More expensive than soaps
- b) Less affected by ions causing water hardness
- c) Naturally occurring materials
- d) Used only with water free of  $Mg^{2+}$  and  $Ca^{2+}$

→ cause water hardness

52) Hardness of water is due to the presence of certain

- a) Alkali metal cations
- b) Transition metal cations
- c) Halogens
- d) None of the above

right answer = alkaline earth metals

53) The impact of household cleaner on the environment is minimized because these products are formulated:

- a) To work with water and end up in wastewater treatment plants
- b) Free of chemicals
- c) With natural ingredients only
- d) To change into environmentally friendly ingredients

54) In hair coloring the oxidation of natural hair pigments to colorless products is achieved by using a

- a) dye
- b) reducing agent
- c) bleaching agent
- d) conditioning agent

55) The oxidizing agent generally used in hair coloring is

- a)  $H_2O$
- b)  $H_2O_2$
- c)  $O_3$
- d)  $H_3O^+$

hydrogen peroxide

56) Sunscreen creams are formulated with additional ingredients to

- a) Charge a higher price
- b) Absorb odors
- c) Make them look colourful
- d) None of the above

right  
answer: absorb UV radiation

57) Fragrances in perfumes are generally experienced in three stages or notes because of the

- a) color of the ingredients
- b) density of the ingredients
- c) price of the ingredients
- d) none of the above

right answer:  
molecular size & volatility

58) Fragrances for perfumes are created by using

- a) Naturally occurring compounds only
- b) Synthetic compounds only
- c) Both natural and synthetic compounds
- d) None of the above

59) The effective chemical ingredients in deodorants are antibacterial agents to

- a) Culture bacteria
- b) Destroy bacteria responsible for converting certain compounds to unpleasant odors
- c) Mask the unpleasant odors
- d) Trap the unpleasant odors

60) Curling of hair involves

- a) Treating hair with a reducing agent
- b) Setting hair in the desired shape
- c) Treating hair with an oxidizing agent
- d) All of the above

61) Lipstick and mascara are examples of

- a) Cosmeceuticals → plastic surgery
- b) Sunscreens
- c) Facial cosmetics → covers stuff up
- d) Pharmaceuticals

62) Carbon-carbon double bond in organic compounds is formed by the sharing of

- a) one pair of electrons
- b) two pairs of electrons
- c) loss and gain of electrons
- d) three pairs of electrons

63) Long chain and cyclic structures are common for organic compounds because carbon can form

- a) ionic bonds with other carbon atoms
- b) covalent bonds with other carbon atoms
- c) no bonds with other carbon atoms
- d) none of the above

C = non metal

2 non metals = covalent

64) The carbon-carbon single bond in organic compounds is

- a) Ionic
- b) Covalent
- c) Carbonic
- d) None of the above

65) The functional group in amine is

- a) -SH
- b) -COOH
- c) -OH
- d) None of the above

right answer.



66) The compound with the formula  $\text{CH}_3\text{CH}_2\text{OCH}_2$  is an

- a) Ester
- b) Ether
- c) Alcohol
- d) None of the above

↓  
makes it  
ether

67) The compound with the formula  $\text{CH}_2(\text{CH}_2)_3\text{OH}$  is an

- a) Amine
- b) Ether
- c) Ester
- d) Alcohol

↓  
alcohol

68) How many carbon atoms are there in a compound with the formula  $\text{CH}_2(\text{CH}_2)_4\text{C}_4\text{H}_2$

- a) 8
- b) 9
- c) 6
- d) 20

69) How many hydrogen atoms are there in a compound with the formula  $\text{CH}_3(\text{CH}_2)_6\text{CH}_2\text{OH}$ ?

- a) 4
- b) 18
- c) 8
- d) 10

70) How many oxygen atoms are there in a compound with the formula  $\text{HOCH}_2(\text{CH}_2)_2\text{COOC}_3\text{H}_7$ ?

- a) 7
- b) 14
- c) 10
- d) 3

71) Benzene has a ring structure containing three carbon-carbon double bonds

- a) True
- b) False

also 3, single bonds

72) The number of carbon atoms in a cyclohexane ring is

- a) 3
- b) 9
- c) 12
- d) None of the above

right answer = 6

H atoms = 12

73) Chemotherapy involves treatment of cancer using

- a) Hypnotherapy
- b) Acupuncture
- c) Alkylating agents
- d) None of the above

74) In the hormone therapy for the treatment of cancer, the chemotherapeutic agent acts against the topoisomerase enzyme to prevent cancer cell growth

- a) true
- b) False

75) The treatment of depression involves the use of

- a) Penicillins
- b) Anti-inflammatory agents
- c) Protease inhibitors
- d) None of the above

SSRI's = right answer

76) The treatment of AIDS involves the use of

- a) Antibacterial agents
- b) OTC's
- c) Antiviral agents
- d) Antidepressants

77) Certain steroids are used as anti-inflammatory drugs because these drugs are

- a) Potent steroids
- b) Capable of suppressing the inflammatory responses of the immune system
- c) Antidepressants
- d) Radioactive

78) The agreement to reduce and ultimately ban the manufacture and use of CFC's goes by the name of

- a) Kyoto Protocol → global warming
- b) Glen Eagle Agreement
- c) Montreal Protocol → CFC's
- d) Geneva Convention

79) The weakly acidic nature of rain is due to the formation of

- a)  $O_2$
- b)  $N_2$
- c)  $H_2CO_3$  → causes acid
- d)  $O_3$

80) the pH of a water sample is measured to be 1.9. The sample is

- a) Acidic
- b) Neutral
- c) Basic
- d) None of the above

81) The pH value for an aqueous solution of citric acid should be

- a) Higher than 7.0
- b) Equal to 11.0
- c) Equal to 7.0
- d) None of the above

right answer:  
less than 7.0

- 82) In the stratosphere, oxygen and ozone interact with solar energy to form a
- a) Dynamic system
  - b) Static system
  - c) Solar system
  - d) No System
- 83) The solubility of lead at a pH of 7.0 is lower than at a pH of 1.2
- a) True
  - b) False
- lower pH = higher solubility*
- 84) Adverse effects of acid rain on the aquatic life are due to
- a) Increase in pH
  - b) Decrease in pH
  - c) Decrease in aqueous concentrations of Al
  - d) Decrease in aqueous concentrations of Pb
- 85) The reaction of limestone with  $\text{H}_2\text{SO}_4$  (aq) generates gaseous
- a) Carbon dioxide
  - b) Sulfur dioxide
  - c) Carbon monoxide
  - d) Sulfur trioxide
- limestone + Acid = carbon dioxide  
(anything beginning with H)*
- 86) Solar energy is an example of a
- a) Fossil energy source
  - b) Non-fossil energy source
  - c) No energy source
  - d) Nuclear energy source
- Solar, wind, nuclear*
- 87) One of the gases responsible for global warming is
- a) CO
  - b)  $\text{N}_2\text{O}$
  - c)  $\text{N}_2$
  - d)  $\text{O}_2$
- $\text{N}_2\text{O}$ ,  $\text{CH}_4$  (methane),  $\text{CO}_2$*
- 88) Which of the following is a greenhouse gas?
- a) Nitrogen
  - b) Oxygen
  - c) Methane
  - d) None of the above

89) One suggested way to control global warming is

- opposite*
- a) Increase CO<sub>2</sub> emissions
  - b) Use more fossil fuels
  - c) Destroy CO<sub>2</sub> sinks
  - d) Conserve energy

90) Welders are at risk of being exposed to ozone in their profession

- a) True
- b) False

91) The increased intensity of ground-level ultraviolet radiation due to ozone depletion may lead to higher probability (in case of humans) of

- a) skin cancer
- b) deafness
- c) depression
- d) none of the above

92) Ground level O<sub>3</sub> is produced during photochemical reactions between

- a) VOCs and O<sub>2</sub>
- b) H<sub>2</sub>O and O<sub>2</sub>
- c) VOCs and NO<sub>x</sub> in the presence of sunlight
- d) VOCs and N<sub>2</sub> in the presence of O<sub>2</sub>

93) Serious damage to lungs may occur on exposure to

- a) O<sub>2</sub>
- b) NO<sub>2</sub>
- c) N<sub>2</sub>
- d) CO<sub>2</sub>

94) Reye syndrome is a possible side effect to the consumption of

- a) Trans fats
- b) Candies
- c) Ice Creams
- d) None of the above

*caused by aspirin → causes brain damage in kids*  
*right answer*

95) It is believed that the effectiveness of Aspirin is due to its ability to lower the production of prostaglandins

- a) True
- b) False

96) The criterion generally used to designate a drug with OTC status involves

- a) Price
- b) Availability of the active ingredient
- c) Benefit – risk comparison
- d) None of the above

↓  
Over-the-counter

97) When Botox is used in providing a more youthful look it is classified as

- a) An OTC drug
- b) A Cosmeceutical
- c) An anti-inflammatory drug
- d) An antiviral drug

98) Alkalating agents are used as chemotherapeutic agents.

- a) True
- b) False

↓ fight cancer by using chemicals

99) Plant debris rest on surface of water bodies due to the

- a) low surface tension of water
- b) no surface tension of water
- c) taste of water
- d) high surface tension of water

Water has highest surface tension of all liquids

100) Temperature extremes affect the working efficiency as well as the health of the worker

- a) true
- b) false