

## Thermodynamics

↳ relationship of heat and temperature to energy and work

Temperature and heat are not the same thing.

Temperature? → What does it do?

- generally it causes them to expand
- in a closed volume of gas, it increases the pressure
- changes electrical resistance of conducting wire

Temperature is defined by zeroth law of thermodynamics

# Conversion between T scales

3-2

| <u>Water</u> | freeze | Boil  | Diff. |
|--------------|--------|-------|-------|
|              | 0°C    | 100°C | 100°C |
|              | 32°F   | 212°F | 180°F |

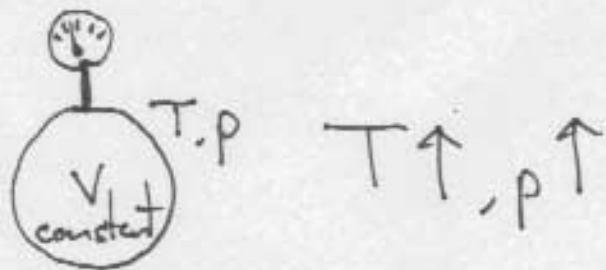
$$T_F(T_c) = \frac{180^\circ\text{F}}{100^\circ\text{C}} (T_c) + \text{const.}$$

$$T_F(0) = \frac{9}{5}(0) + \text{const} = 32$$

const = 32

$$T_F = \frac{9}{5}T_c + 32^\circ\text{C}$$

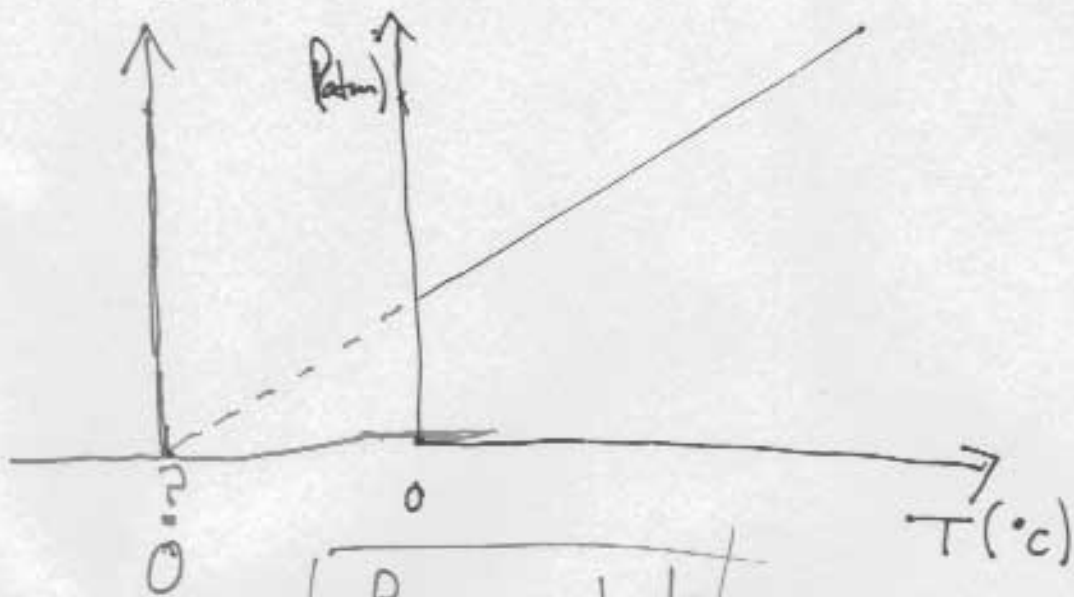
17.3 Ideally, our T scale will not depend on the properties of a particular material.



Many gases (ideal)

3-3

$$\frac{\Delta P}{\Delta T} = \text{constant}$$



$$\frac{P}{T} = \text{constant}$$

Meaning that  $\frac{P_1}{T_1} = \frac{P_2}{T_2}$

or

$$\frac{P_1}{P_2} = \frac{T_1}{T_2} \quad \begin{array}{l} 1. \text{ constant } V \\ 2. T \text{ in } K \end{array}$$

Kelvin T scale.

## Example

3-4

A constant volume gas thermometer is calibrated in dry ice ( $\text{CO}_2$  in solid state,  $-80.0^\circ\text{C}$ ) and in boiling ethyl alcohol ( $78.0^\circ\text{C}$ ). The pressure readings are  $0.900 \text{ atm}$  and  $1.635 \text{ atm}$ . What is the temperature value in  $^\circ\text{C}$  when extrapolating back to  $p=0$ ?

LINEAR:  $P = A + BT$   $A, B$  are constants

$$\textcircled{1} 0.900 \text{ atm} = A + (-80.0^\circ\text{C})B$$

$$\textcircled{2} 1.635 \text{ atm} = A + (78.0^\circ\text{C})B$$

$$A = 0.900 \text{ atm} + (80.0^\circ\text{C})B$$

$\hookrightarrow \textcircled{2}$

$$B = 4.65 \times 10^{-3} \frac{\text{atm}}{^\circ\text{C}}$$

$$A = 1.27 \text{ atm}$$

$$P = 1.27 \text{ atm} + (4.65 \times 10^{-3} \text{ atm}/^\circ\text{C})T$$

$$P = 0 = 1.27 \text{ atm} + (4.65 \times 10^{-3} \text{ atm}/^\circ\text{C})T$$

solve for  $T$

$$T = -274^\circ\text{C}$$

Absolute zero:  $0\text{ K}$ ,  $-273.15^\circ\text{C}$

because all we did was shift our scale

$$T_{\text{K}} = T_{\text{C}} + 273.15^\circ\text{C}$$

#### 17.4 | Thermal Expansion

## Example

3-6

The active element of a laser is made of a glass rod 30.0 cm long by 1.50 cm in diameter. If the temperature of the rod increases by 65.0°C, what is the increase in

a) its length?

b) its diameter?

$$\begin{aligned} \text{a) } \Delta L &= \alpha L_0 \Delta T \\ &= (9.00 \times 10^{-6} \text{ } ^\circ\text{C}^{-1})(0.300 \text{ m})(65.0 \text{ } ^\circ\text{C}) \\ &= 1.76 \times 10^{-4} \text{ m} \end{aligned}$$

b) Diameter is a linear dimension, so same equation applies

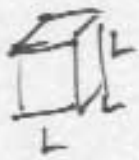
$$\begin{aligned} \Delta D &= \alpha D_0 \Delta T \\ &= (9.00 \times 10^{-6} \text{ } ^\circ\text{C}^{-1})(0.0150 \text{ m})(65.0 \text{ } ^\circ\text{C}) \\ &= 8.78 \times 10^{-6} \text{ m} \end{aligned}$$

What about volume?

$\beta$  - coefficient of volume expansion

$$\Delta V = \beta V_0 \Delta T$$

Can we relate  $\alpha$  to  $\beta$ ?



$$V_0 = L^3$$

3-7

change in volume  $dV = \frac{dV}{dL} dL = 3L^2 dL$

$$dL = \alpha L_0 dT$$

$$V_0 = L_0^3$$

$$dV = 3L_0^2 \alpha L_0 dT$$

$$= \underbrace{3\alpha}_{\beta} V_0 dT$$

Increase in volume for glass rod?

$$V_0 = \frac{\pi}{4} r^2 L = \frac{\pi}{4} (0.0150 \text{ m})^2 (0.300 \text{ m})$$
$$= 5.30 \times 10^{-5} \text{ m}^3$$

$$\Delta V = \beta V_0 \Delta T = 3\alpha V_0 \Delta T$$
$$= 93.0 \times 10^{-9} \text{ m}^3$$

A hollow aluminum ( $\beta = 7.2 \times 10^{-5} \text{ }^\circ\text{C}^{-1}$ ) cylinder 20.0 cm deep has an internal capacity of 2.000 L at 20.0°C. It's completely with turpentine, and then slowly warm to 80.0°C.

3-8

a) How much turpentine overflows?

b) If the cylinder is then cooled back to 20.0°C, how far below the cylinder's rim does the surface of the turpentine recede?

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a)  $\Delta V = \beta V_0 \Delta T$

The overflow

$$V_{\text{OVER}} = \Delta V_T - \Delta V_{AL}$$

$$= (\beta_T V_0 \Delta T) - (\beta_{AL} V_0 \Delta T)$$

$$= V_0 \Delta T (\beta_T - \beta_{AL})$$

$$= (2.000 \text{ L})(60.0 \text{ }^\circ\text{C})(9.00 \times 10^{-4} \text{ }^\circ\text{C}^{-1} - 0.72 \times 10^{-4} \text{ }^\circ\text{C}^{-1})$$

$$= 0.0994 \text{ L}$$

b) After warming the whole volume of turpentine is <sup>3-9</sup>

$$V' = V_0 + \beta_T V_0 \Delta T$$

$$= 2.000 \text{ L} + (9.00 \times 10^{-4} \text{ } ^\circ\text{C}^{-1})(2.000 \text{ L})(60.0^\circ\text{C})$$

$$= 2.108 \text{ L}$$

Fraction lost is

$$\frac{0.0994 \text{ L}}{2.108 \text{ L}} = 4.7 \times 10^{-2} \quad \text{or } 4.7\%$$

$$\Delta h = (4.71 \times 10^{-2})(20.0 \text{ cm})$$

$$= 0.943 \text{ cm}$$

Thermal Stress

$$\frac{F}{A} = -Y\alpha\Delta T$$

F - tension

A - x sectional area

Y - Young's Modulus - ratio of stress to strain in a material

$\alpha$  - coef. of linear expansion

$\Delta T$  - T change

$\Delta T \downarrow$ ,  $\Delta T$  is negative,  $F$  and  $\frac{F}{A}$  are ~~negative~~  
positive

$\hookrightarrow$  a tensile force is needed to  
maintain the length of the material

$\Delta T \uparrow$ ,  $F$  and  $\frac{F}{A}$  are negative

$\hookrightarrow$  force is compressive