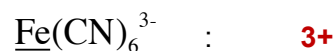


## # 1. (10 points) Short Answer Questions

a) Indicate the oxidation state of the underlined atom: (2 pts)



b) Provide the name or formula for the following: (2 pts)

lead (II) sulfate  $\text{PbSO}_4$  $\text{NH}_4\text{Cl}$  ammonium chloride

c) Give the precise chemical symbol for the following: (1 pt)

Number of protons	Number of neutrons	Number of electrons	Chemical Symbol
35	46	36	${}_{35}^{81}\text{Br}^-$

d) The standard heat of formation of solid magnesium sulfate is  $-1278.2$  kJ. Write the chemical equation for the reaction to which this value applies (include phases). (1 pt)e) Number of atoms in 0.195 mol of  $\text{FeCl}_3$ : (1 pt)  $4.70 \times 10^{23}$  atoms

f) Four balloons are filled to the same volume with one of the following gases. If a small hole is made in each balloon, which one will deplete faster? (circle one) (1 pt)

 $\text{NO}_2$ 

Xe

Kr

 $\text{SH}_2$ 

g) You take a 2.0 L volume of gas at 250 K. You compress the gas to a 1.0 L volume and simultaneously heat the vessel to 500 K. The pressure of the gas: (circle one) (1 pt)

DOUBLES

HALVES

QUADRUPLES

STAYS CONSTANT

h) How much heat (in kJ) must be transferred to a 1.0 L of water in a pot to raise its temperature from  $25^\circ\text{C}$  to  $90^\circ\text{C}$ : (1 pt)  $+ 272$  kJ ( $q = ms\Delta T$ )

# 2. (10 points) A parachute manufacturer decides to start making its own adipic acid ( $\text{H}_2\text{C}_6\text{H}_8\text{O}_4$ ), one of the main components of nylon polymers. It uses a controlled reaction between cyclohexane ( $\text{C}_6\text{H}_{12}$ ) and  $\text{O}_2$  shown here (not balanced):



$$\text{MM} = \quad 84.16 \text{ g/mol} \quad 32.00 \text{ g/mol} \quad 146.1 \text{ g/mol}$$

The reaction uses 29.0 g of cyclohexane and 31.0 g of oxygen, and allows the isolation of 45.2 g of adipic acid. What are the theoretical yield of adipic acid and the % yield of the reaction?

### **DETERMINING THE LIMITING REAGENT:**

$29.0 \text{ g C}_6\text{H}_{12} = 0.345 \text{ mol}$ $31.0 \text{ g O}_2 = 0.969 \text{ mol}$ $\text{Ideal mol ratio: } = \frac{\text{O}_2}{\text{C}_6\text{H}_{12}} = \frac{5 \text{ O}_2}{2 \text{ C}_6\text{H}_{12}} = 2.5$ $\text{Actual mol ratio: } = \frac{\text{O}_2}{\text{C}_6\text{H}_{12}} = \frac{0.969 \text{ mol}}{0.345 \text{ mol}} = 2.81 > 2.5$ <p>Hence: <math>\text{C}_6\text{H}_{12}</math> is limiting (small excess of <math>\text{O}_2</math>)</p>
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### **THEORETICAL YIELD:**

$? \text{ g acid} = 29.0 \text{ g C}_6\text{H}_{12} \cdot \frac{\text{mol C}_6\text{H}_{12}}{84.16 \text{ g C}_6\text{H}_{12}} \cdot \frac{1 \text{ mol acid}}{1 \text{ mol C}_6\text{H}_{12}} \cdot \frac{146.1 \text{ g acid}}{\text{mol acid}}$ $= 50.3 \text{ g adipic acid}$
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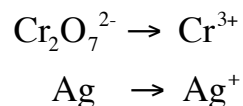
### **ACTUAL YIELD:**

$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \cdot 100\% = \frac{45.2 \text{ g}}{50.3 \text{ g}} \cdot 100\%$ $= 89.9 \%$
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Theoretical Yield: **50.3 g**

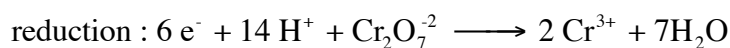
% Yield: **89.9 %**

# 3. (10 points) In photography, black and white transparencies (black pictures on a transparent background) are produced from the oxidation (*bleaching*) of the metallic silver covering the film to silver ions. Dichromate ions are commonly used as oxidizing agents, in an acidic solution. The following half reactions are involved in the redox process:

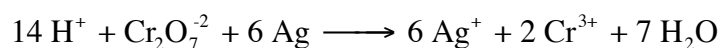


a) Determine the overall balanced redox reaction equation. (6 pts)

**HALF-REACTIONS AND BALANCING OF ELECTRONS:**



**OVERALL BALANCED REDOX:**



b) A black and white negative film is covered with 5.55 g of metallic silver. What is the minimum volume (in mL) of a 0.250 M  $\text{K}_2\text{Cr}_2\text{O}_7$  that must be added to the bleaching bath to ensure a complete oxidation of the metallic silver present? (4 pts)

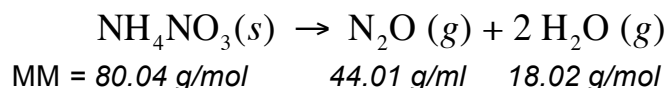
**SOLUTION:**

$$\begin{aligned} ? \text{ mL of } \text{K}_2\text{Cr}_2\text{O}_7 &= 5.55 \text{ g Ag} \cdot \frac{\text{mol Ag}}{107.8 \text{ g Ag}} \cdot \frac{1 \text{ mol } \text{K}_2\text{Cr}_2\text{O}_7}{6 \text{ mol Ag}} \cdot \frac{1 \text{ L}}{0.250 \text{ mol } \text{K}_2\text{Cr}_2\text{O}_7} \cdot \frac{1000 \text{ mL}}{\text{L}} \\ &= 34.3 \text{ mL of } \text{K}_2\text{Cr}_2\text{O}_7 \text{ solution} \end{aligned}$$

Answer:

**34.3 mL**

# 4. (10 points) A 4.30 g sample of ammonium nitrate is introduced into an evacuated 2.15 L flask and then sealed and heated to 250 °C, causing the following decomposition:



a) What is the total gas pressure *in mmHg* in the flask once the  $\text{NH}_4\text{NO}_3(s)$  has completely decomposed (assume a complete conversion to the products)? (5 pts)

$$? \text{ mol gas formed} = 4.30 \text{ g NH}_4\text{NO}_3 \cdot \frac{\text{mol NH}_4\text{NO}_3}{80.04 \text{ g NH}_4\text{NO}_3} \cdot \frac{3 \text{ mol of gas formed}}{1 \text{ mol NH}_4\text{NO}_3} = 0.1612 \text{ mol of gas}$$

$$\text{since } PV = nRT, \quad P = \frac{nRT}{V} = \frac{0.1612 \text{ mol} \cdot 0.08206 \frac{\text{atm L}}{\text{K mol}} \cdot 523 \text{ K}}{2.15 \text{ L}} = 3.22 \text{ atm}$$

$$P(\text{in mmHg}) = 3.22 \text{ atm} \cdot \frac{760 \text{ mmHg}}{\text{atm}} = 2447 \text{ mmHg}$$

Answer:      **2.45 x 10<sup>3</sup> mmHg**

b) Pure  $\text{N}_2\text{O}(g)$  (also known as nitrous oxide or laughing gas) can be used in a mixture with  $\text{O}_2(g)$  as an anesthetic. In a particular mixture of this type, the partial pressures of  $\text{N}_2\text{O}(g)$  and  $\text{O}_2(g)$  are 600 mmHg and 166 mmHg respectively. Calculate the mass percentage of  $\text{O}_2$  in the mixture. (MM  $\text{O}_2$  = 32.00 g/mol) (5 pts)

**MOLAR FRACTIONS:**

$$\chi_{\text{N}_2\text{O}} = \frac{P_{\text{N}_2\text{O}}}{P_{\text{total}}} = \frac{600 \text{ mmHg}}{760 \text{ mmHg}} = 0.783$$

$$\chi_{\text{O}_2} = \frac{P_{\text{O}_2}}{P_{\text{total}}} = \frac{166 \text{ mmHg}}{760 \text{ mmHg}} = 0.217$$

**MOLAR MASS FOR 1 MOL OF GAS:**

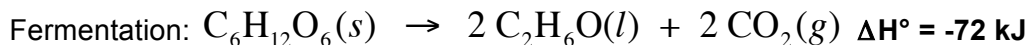
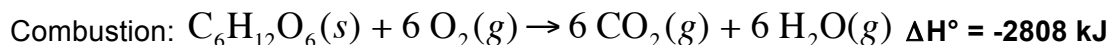
$$MM = 0.783 \cdot 44.01 \text{ g/mol} + 0.217 \cdot 32.00 \text{ g/mol} = 41.40 \text{ g/mol}$$

**MOLAR MASS FOR 1 MOL OF GAS:**

$$\% \text{ O}_2 = \frac{\text{mass of O}_2 \cdot 100\%}{\text{MM of gas}} = \frac{\chi_{\text{O}_2} \cdot \text{MM of O}_2 \cdot 100\%}{\text{MM of gas}} = \frac{0.217 \cdot 32.00 \text{ g/mol} \cdot 100\%}{41.40 \text{ g/mol}} = 16.8 \%$$

Answer:      **16.8 %**

# 5. (10 points) The standard enthalpies for the combustion of glucose ( $C_6H_{12}O_6$ ) and for the fermentation of glucose to ethanol ( $C_2H_6O$ ) are shown below:



a) How much work is performed by the system during the combustion of glucose at 25 °C? (3 pts)

$$\begin{aligned} \Delta n &= \# \text{ mol gas in products} - \# \text{ mol gas in reactants} \\ &= 12 \text{ mol} - 6 \text{ mol} = 6 \text{ mol} \end{aligned}$$

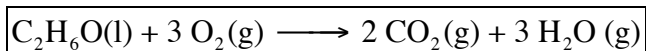
$$\begin{aligned} W_{\text{system}} &= -P\Delta V = -\Delta nRT = -6 \text{ mol} \cdot 8.31451 \text{ J K}^{-1} \text{ mol}^{-1} \cdot 298 \text{ K} \\ &= -14873 \text{ J} = -14.873 \text{ kJ} \end{aligned}$$

Answer: **-14.9 kJ**

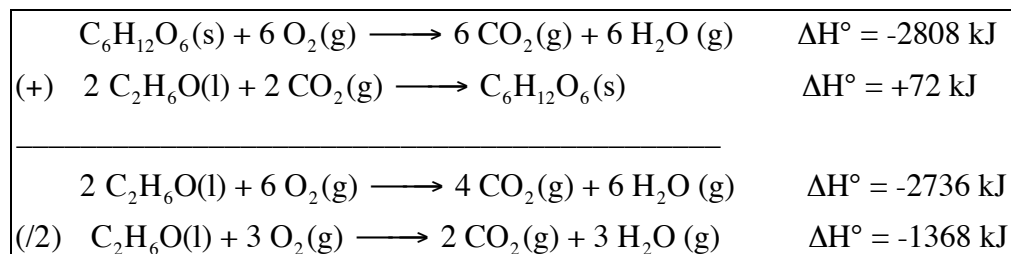
b) Explain briefly the significance of the sign (+ or -) of your answer in a). (2 pts)

(-) negative work is done by the system on the surroundings (expansion of the gas)

c) Provide a balanced equation that shows the combustion of ethanol  $C_2H_6O(l)$ . (2 pts)



d) Use the standard enthalpies of the reactions provided in the question to calculate the standard enthalpy for the combustion of 1 mol of ethanol  $C_2H_6O(l)$ . (3 pts)



Answer: **-1368 kJ**